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(54) Title: A PROCESS AND A CATALYST FOR PREVENTING REACTOR FOULING			
(57) Abstract A support containing methylalumoxane and derivatives thereof is described which is formed by an incipient impregnation technique. The most preferred support is silica. Incipient impregnation in accordance with the invention provides a supported alum xane, methylalumoxane, which substantially eliminates the problem of fluidized bed reactor fouling when methylalumoxane is introduced into the reactor during its operation. In accordance with the invention, the process comprises providing methylalumoxane activated metallocene compound in particulate form as catalysts in fluidized bed gas phase operation.			

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A PROCESS AND A CATALYST FOR PREVENTING REACTOR FOULING

The invention relates to modifications and improvements in low pressure fluidized bed gas phase systems for the polymerization and copolymerization of ethylene, undertaken in the presence of catalysts comprising metallocenes of transition metals. The invention relates to eliminating reactor fouling and to maintaining the continuous operation of the distributor plate in the fluidized bed gas phase reactor undertaken in the presence of catalysts comprising metallocenes of transition metals. Moreover, the invention relates to maintaining the continuous polymerization and copolymerization of ethylene in the fluidized bed reactor.

Metallocenes can be described by the empirical formula $Cp_nMA_nB_p$. These compounds in combination with methylalumoxane (MAO) have been used to produce olefin polymers and copolymers, such as ethylene and propylene homopolymers, ethylene-butene and ethylene-hexene copolymers.

Methylalumoxane (MAO) is used as co-catalyst with metallocene catalysts. The class of alumoxanes comprises oligomeric linear and/or cyclic alkylalumoxanes represented by the formula:

$R-(Al(R)-O)_n-AlR_2$ for oligomeric, linear alumoxanes and $(-Al(R)-O)_n$ for oligomeric cyclic alumoxanes

wherein n is 1-40, preferably 10-20, m is 3-40, preferably 3-20 and R is a C_1-C_8 alkyl group and preferably methyl.

Methylalumoxane is commonly produced by reacting trimethylaluminum with water or with hydrated inorganic salts, such as $Cu(SO_4, 5H_2O)$ or $Al_2(SO_4)_3 \cdot 5H_2O$. Methylalumoxane can be also generated in situ in polymerization reactors by adding to the reactor trimethylaluminum and water or water-containing inorganic salts. MAO is a mixture of oligomers with a very wide distribution of molecular weights and usually with an average molecular weight of 1200. MAO is typically kept in solution in toluene. While the MAO

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solutions remain liquid at fluidized bed reactor temperatures, the MAO itself is a solid at room temperature.

Most of the experiments reported in the literature relating to methylalumoxane used as a cocatalyst with metallocene catalysts are undertaken in a slurry or solution process, rather than in a gas phase fluidized bed reactor process.

It was found that the metallocene compound must contact the MAO cocatalyst while MAO is in solution in order for the metallocene compound to be activated in the fluidized bed reactor. Moreover, it was discovered that extensive reactor fouling results when MAO solutions are fed directly into the gas phase reactor in large enough quantities to provide this liquid contact. The fouling occurs because the MAO solution forms a liquid film on the interior walls of the reactor. The metallocene compound is activated when it comes into contact with this liquid film, and the activated catalyst reacts with ethylene to form a polymer coating which grows larger in size until the reactor is fouled. In addition, since substantially all of the activation takes place on the walls, the MAO is not uniformly distributed to the catalyst particles. The resulting non-homogeneous polymerization gives low catalyst activity and poor product properties.

The problems invoked by the use of an alumoxane, methylalumoxane, in catalyst production are addressed by a process for forming a carrier material impregnated with alumoxane and derivatives thereof, comprising

(1) providing a carrier, which has hydroxyl groups, which is porous and has a particle size of 1 to 200 microns, having pores which have an average diameter of 50 to 500 Angstroms and having a pore volume of 0.5 to 5.0 cc/g of carrier;

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(2) providing a volume of a solution comprising a alumoxane and a solvent therefor, wherein the concentration of alumoxane, expressed as Al weight percent is 5 to 20;

wherein the alumoxane provides aluminum in an amount
5 sufficient to provide a ratio of Al to silica, i.e.,
Al/silica (wt/wt) ranging from 0.10 to 0.40;

(3) contacting the said solution with the carrier and
allowing the solution to impregnate the pores of the
carrier, without forming a slurry of the carrier in said
10 solution having a pore volume of 0.5 to 5.0 cc/g,
containing alumoxane within said pores.

(4) after said contacting, recovering a dry
impregnated carrier.

Accordingly, the invention includes the foregoing
15 process, the product(s) thereof, the catalysts formed
therefrom, and the polymerizations and copolymerizations
achieved thereby.

Ethylene polymers, as well as copolymers of ethylene
with one or more C₃-C₁₀ alpha-olefins, can be produced in
20 accordance with the invention. Thus, copolymers having two
monomeric units are possible as well as terpolymers having
three monomeric units. Particular examples of such
polymers include ethylene/1-butene copolymers, ethylene/1-
hexene copolymers and ethylene/4-methyl-1-pentene
25 copolymers.

Hydrogen may be used as a chain transfer agent in the
polymerization reaction of the present invention. The
ratio of hydrogen/ethylene employed will vary between 0 to
2.0 moles of hydrogen per mole of ethylene in the gas
30 phase. Any other gas inert to the catalyst and reactants
can also be present in the gas stream.

Ethylene/1-butene and ethylene/1-hexene copolymers are
the most preferred copolymers polymerized in the process of
and with the catalyst of this invention. The ethylene
35 copolymers produced in accordance with the present
invention preferably contain at least 70 percent by weight

of ethylene units. The cocatalyst of this invention can also be used with the catalyst precursor of this invention to polymerize propylene and other alpha-olefins and to copolymerize them. The structure of alpha-olefin polymers prepared with the cocatalyst and the catalyst precursor of this invention depends on the structure of the cyclopentadienyl ligands attached to the metal atom in the catalyst precursor molecule. The cocatalyst compositions of this invention can also be used with the catalyst precursors of this invention to polymerize cyclic-olefins such as cyclopentene.

In one embodiment, a catalyst of the invention exhibits high activity for polymerization of ethylene and higher alpha-olefins and allows the synthesis of ethylene polymers and copolymers with a relatively narrow molecular weight distribution and homogeneous branching distribution. The catalyst of the invention exhibits high activity for copolymerization of ethylene and higher alpha-olefins. The molecular weight distribution is determined as MFR which ranges from 15 to 25, in polymerizations of the invention. Branching distribution in ethylene copolymers is evaluated on the basis of the resin's melting point. Relatively homogeneous branching distribution is one in which the melting point of the polymer ranges from 100 to 140°C, depending on comonomer composition. In this embodiment, the catalyst of the invention contains only one source of transition metal, a metallocene compound.

In another embodiment of the invention, a catalyst of the invention exhibits high activity for polymerization of ethylene and higher alpha-olefins and allows the synthesis of ethylene polymers and copolymers with a broad/bimodal molecular weight distribution and generally, bimodal molecular weight distribution with a relatively high molecular weight component and with a relatively lower molecular weight component in the resin. The molecular weight distribution of the broad/bimodal resin, expressed

as MFR, is 50 to 250. In this embodiment, the catalyst of the invention comprises two transition metal compounds, only one of the transition metal compounds being a metallocene.

5 The Fluidized Bed Reactor

A fluidized bed reactor which can be used in the practice of the process of the present invention includes consists of a reaction zone, a velocity reduction zone and the distributor plate. Although fouling can occur in all
10 of the cold areas (areas in a reactor at a temperature which is less than the temperature at which any component(s), in the gas phase reactor are liquid rather than gaseous) distributor plate fouling is the one most easily detected, since it results in a rapid increase in
15 the pressure drop across the distributor plate due to flow restriction. Such flow restrictions also result in changing fluidization patterns and contribute to reactor wall fouling. The lowest temperature in the reactor loop is in the reactor inlet beneath the distributor plate.
20 Other areas representing the coldest sections in the fluidized bed reactor system include the cooler and piping between the cooler and the bottom head.

The reaction zone comprises a bed of growing polymer particles and a minor amount of catalyst particles
25 fluidized by the continuous flow of polymerizable and modifying gaseous components. To maintain a viable fluidized bed, the mass gas flow rate through the bed must be above the minimum flow required for fluidization, and preferably from 1.5 to 10 times G_{mf} and more preferably from
30 3 to 6 times G_{mf} . G_{mf} is used in the accepted form as the abbreviation for the minimum mass gas flow required to achieve fluidization, C. Y. Wen and Y. H. Yu, "Mechanics of Fluidization", Chemical Engineering Progress Symposium Series, Vol. 62, p. 100-111 (1966). Th distribution plat
35 serv s the purpose of diffusing r cycle gas through the bed at a rate suffici nt to maintain fluidization at the bas

of the bed. Fluidization is achieved by a high rate of gas recycle to and through the bed, typically in the order of 50 times the rate of feed of make-up gas. Make-up gas is fed to the bed at a rate equal to the rate at which
5 particulate polymer product is formed by reaction. The composition of the make-up gas is determined by a gas analyzer positioned above the bed. The composition of the make-up gas is continuously adjusted to maintain an essentially steady state gaseous composition within the
10 reaction zone.

The portion of the gas stream which does not react in the bed (the recycle gas) passes a velocity reduction zone and is compressed in a compressor, passes through a heat exchanger and is returned to the bed. The distribution
15 plate 20 serves the purpose of diffusing recycle gas through the bed at a rate sufficient to maintain fluidization. The plate may be a screen, slotted plate, perforated plate, a plate of the bubble cap type, and the like. The elements of the plate may all be stationary, or
20 the plate may be of the mobile type disclosed in U.S. Patent No. 3,298,792.

Conditions in the fluidized bed reactor for the gas phase polymerization and copolymerization of ethylene

It is essential to operate the fluidized bed reactor
25 at a temperature below the sintering temperature of the polymer particles. For the production of ethylene copolymers in the process of the present invention an operating temperature of 30° to 115°C is preferred, and a temperature of 75° to 95°C is most preferred. Temperatures
30 of 75° to 90°C are used to prepare products having a density of 0.91 to 0.92, and temperatures of 80° to 100°C are used to prepare products having a density of 0.92 to 0.94, and temperatures of 90° to 115°C are used to prepare products having a density of 0.94 to 0.96.

35 The fluidized bed reactor is operated at pressures of up to 1000 psi, and is preferably operated at a pressure of

from 150 to 350 psi, with operation at the higher pressures in such ranges favoring heat transfer since an increase in pressure increases the unit volume heat capacity of the gas.

5 The partially or completely activated catalyst is injected into the bed at a point above the distribution plate at a rate equal to its consumption. Since the catalysts used in the practice of this invention are highly active, injection of the fully activated catalyst into the
10 area below the distribution plate may cause polymerization to begin there and eventually cause plugging of the distribution plate. Injection into the bed, instead, aids in distributing the catalyst throughout the bed and precludes the formation of localized spots of high catalyst
15 concentration.

 The production rate of polymer in the bed is controlled by the rate of catalyst injection. Since any change in the rate of catalyst injection changes the rate of generation of the heat of reaction, the temperature of
20 the recycle gas is adjusted to accommodate the change in rate of heat generation. Complete instrumentation of both the fluidized bed and the recycle gas cooling system is, of course, necessary to detect any temperature change in the bed so as to enable the operator to make a suitable
25 adjustment in the temperature of the recycle gas.

 Since the rate of heat generation is directly related to product formation, a measurement of the temperature rise of the gas across the reactor (the difference between inlet gas temperature and exit gas temperature) is determinative
30 of the rate of particulate polymer formation at a constant gas velocity.

 Under a given set of operating conditions, the fluidized bed is maintained at essentially a constant height by withdrawing a portion of the bed as product at a
35 rate equal to the rate of formation of the particulate polymer product.

Catalyst Composition

Catalysts which contain only one transition metal in the form of a metallocene have an activity of at least 1,000 g polymer/g catalyst or at least 250 kg polymer/g transition metal. Catalysts which contain two transition metals, one in the form of a metallocene and one transition metal in the form of a non-metallocene, have an activity of at least 1,000 g polymer/g catalyst or 100 kg polymer/g of each transition metal.

The bimetallic catalysts of the invention comprise, optionally, a cocatalyst comprising an aluminum alkyl compound, such as a trialkyl aluminum free of alumoxane, and a catalyst precursor comprising a carrier, an alumoxane and at least one metallocene; in one embodiment the catalysts further include a non-metallocene transition metal compound.

The carrier material is a solid, particulate, porous, preferably inorganic material, such as an oxide of silicon and/or of aluminum and/or of Mg. The carrier material is used in the form of a dry powder having an average particle size of from 1 micron to 250 microns, preferably from 10 microns to 150 microns. If necessary, the treated carrier material may be sieved to insure that the particles of the ultimate carrier-catalyst containing composition are no larger than 250 microns. This is highly desirable in the embodiment of the invention, in which the catalyst contains only one transition metal in the form of a metallocene and which is used to form narrow molecular weight distribution LLDPE, to reduce gels. The surface area of the carrier is at least 3 square meters per gram (m^2/g), and preferably at least 50 m^2/g up to 350 m^2/g . The carrier material should be dry, that is, free of absorbed water. Drying of the carrier material can be effected by heating at 100°C to 1000°C, preferably at 200°C-600°C. When the carrier is silica, it is heated to at least 200°C, preferably 250°C for catalysts containing only a metallocene compound and

600°C for catalysts containing a metallocene compound and a non-metallocene transition metal compound. The carrier material must have at least some active hydroxyl (OH) groups to produce the catalyst composition of this invention.

In the most preferred embodiment, the carrier is silica which, prior to the use thereof in the first catalyst synthesis step, has been dehydrated by fluidizing it with nitrogen and heating at 250°C-600°C for 16 hours to achieve a surface hydroxyl group concentration of at least 0.7 millimoles per gram (mmol/g). The silica of the most preferred embodiment is a high surface area, amorphous silica (surface area = 300 m²/g; pore volume of 1.65 cm³/g), and it is a material marketed under the tradenames of Davison 952 or Davison 955 by the Davison Chemical Division of W.R. Grace and Company. The silica is in the form of spherical particles, e.g., as obtained by a spray-drying process.

To form catalysts of the invention, some catalyst precursor components can be dissolved with alumoxane and impregnated into the carrier. In a unique process, the carrier material is impregnated with a solid alumoxane, preferably methylalumoxane, in a process described below. The class of alumoxanes comprises oligomeric linear and/or cyclic alkylalumoxanes represented by the formula:

$R-(Al(R)-O)_n-AlR_2$ for oligomeric, linear alumoxanes and $(-Al(R)-O-)_m$ for oligomeric cyclic alumoxane

wherein n is 1-40, preferably 10-20, m is 3-40, preferably 3-20 and R is a C₁-C₈ alkyl group and preferably methyl.

MAO is a mixture of oligomers with a very wide distribution of molecular weights and usually with an average molecular weight of 1200. MAO is typically kept in solution in toluene. While the MAO solutions remain liquid at fluidized bed reactor temperatures, the MAO itself is a solid.

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Although the alumoxane can be impregnated into the carrier at any stage of the process of catalyst preparation, the preferred stage of alumoxane incorporation will depend on the ultimate catalyst sought to be synthesized. In a preferred embodiment, of alumoxane incorporation into the carrier, one of the controlling factors in the alumoxane incorporation into the carrier material catalyst synthesis is the pore volume of the silica. In this preferred embodiment, the process of impregnating the carrier material is by infusion of the alumoxane solution, without forming a slurry of the carrier material, such as silica, in the alumoxane solution. The volume of the solution of the alumoxane is sufficient to fill the pores of the carrier material without forming a slurry of the solution and the silica. To do this preferably, the maximum volume of the alumoxane solution is, does not exceed, the total pore volume of the carrier material sample. However, it has been determined that no slurry will form when the volume of the solution is 200% of the total pore volume of the sample of silica. Accordingly, the volume of the alumoxane can range from less than the total pore volume up to twice the total pore volume of the silica, without a silica-alumoxane slurry forming. However, it is preferred to use as a maximum volume of alumoxane solution that is equal to; or less than, the total pore volume of silica. We have found that as the volume of the solution exceeds the total pore volume of the carrier sample, formation, and increase of that formation, of a crust of catalyst on the outside of particle occurs; this layer is detectable by microscopy. As the amount of alumoxane (e.g. methylalumoxane) layer on the particle increases, e.g. zirconium efficiency (activity and productivity) decreases. Although there is no slurry formation, when the volume of the impregnant alumoxane solution employed is up to 200% of the total pore volume of the carrier, preferably the volume is less than 200% of the

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total pore volume of the carrier, more preferably the volume of that solution is less than 150% of the total pore volume of the carrier. That maximum volume of the alumoxane solution insures that no slurry of silica is formed. Accordingly, if the pore volume of the carrier material is $1.65\text{cm}^3/\text{g}$, then the volume of alumoxane will be equal to or less than $1.65\text{ cm}^3/\text{g}$ of carrier material. As a result of this proviso, the impregnated carrier material will appear dry immediately following impregnation although the pores of the carrier will be filled with inter alia solvent. Solvent may be removed from the alumoxane impregnated pores of the carrier material by heating and/or under a positive pressure induced by an inert gas, such as nitrogen. If employed, the conditions in this step are controlled to reduce, if not to eliminate, agglomeration of impregnated carrier particles and/or crosslinking of the alumoxane. In this step, solvent can be removed by evaporation effected at relatively low elevated temperatures of above 40° and below 50° to obviate agglomeration of catalyst particles and crosslinking of the alumoxane. Although solvent can be removed by evaporation at relatively higher temperatures than that defined by the range above 40° and below 50°C , very short heating times schedules must be employed to obviate agglomeration of catalyst particles, crosslinking of the alumoxane, or deactivation of the resulting catalyst precursor.

In a preferred embodiment, the metallocene is added to the solution of the alumoxane prior to impregnating the carrier with the solution. Again the preferred maximum volume of the alumoxane solution also including the metallocene is the total pore volume of the carrier material sample. The mole ratio of alumoxane provided aluminum, expressed as Al, to metallocene metal expressed as M (e.g. Zr), ranges from 50 to 500, preferably 75 to 300, and most preferably 100 to 200. An added advantage of the present invention is that this Al:Zr ratio can be

directly controlled. In a preferred embodiment the alumoxane and metallocene compound are mixed together at a temperature of 20 to 80°C, for 0.1 to 6.0 hours, prior to use in the infusion step. The solvent for the metallocene and alumoxane can be appropriate solvents, such as aromatic hydrocarbons, halogenated aromatic hydrocarbons, ethers, cyclic ethers or esters, preferably it is toluene.

The metallocene compound has the formula $Cp_mMA_nB_p$ in which Cp is an unsubstituted or substituted cyclopentadienyl group, M is zirconium or hafnium and A and B belong to the group including a halogen atom, hydrogen or an alkyl group. In the above formula of the metallocene compound, the preferred transition metal atom M is zirconium. In the above formula of the metallocene compound, the Cp group is an unsubstituted, a mono- or a polysubstituted cyclopentadienyl group. The substituents on the cyclopentadienyl group can be preferably straight-chain C_1 - C_6 alkyl groups. The cyclopentadienyl group can be also a part of a bicyclic or a tricyclic moiety such as indenyl, tetrahydroindenyl, fluorenyl or a partially hydrogenated fluorenyl group, as well as a part of a substituted bicyclic or tricyclic moiety. In the case when m in the above formula of the metallocene compound is equal to 2, the cyclopentadienyl groups can be also bridged by polymethylene or dialkylsilane groups, such as $-CH_2-$, $-CH_2-CH_2-$, $-CR'R''-$ and $-CR'R''-CR'R''-$ where R' and R'' are short alkyl groups or hydrogen, $-Si(CH_3)_2-$, $Si(CH_3)_2-CH_2-CH_2-Si(CH_3)_2-$ and similar bridge groups. If the A and B substituents in the above formula of the metallocene compound are halogen atoms, they belong to the group of fluorine, chlorine, bromine or iodine. If the substituents A and B in the above formula of the metallocene compound are alkyl groups, they are preferably straight-chain or branch d C_1 - C_6 alkyl groups, such as methyl, ethyl, n-propyl, isopropyl, n-butyl, isobutyl, n-pentyl, n-hexyl or n-octyl.

Suitable metallocene compounds include
bis(cyclopentadienyl)metal dihalides,
bis(cyclopentadienyl)metal hydridohalides,
bis(cyclopentadienyl)metal monoalkyl monohalides,
5 bis(cyclopentadienyl)metal dialkyls and bis(indenyl)metal
dihalides wherein the metal is zirconium or hafnium, halide
groups are preferably chlorine and the alkyl groups are C₁-
C₆ alkyls. Illustrative, but non-limiting examples of
metallocenes include bis(cyclopentadienyl)zirconium
10 dichloride, bis(cyclopentadienyl)hafnium dichloride,
bis(cyclopentadienyl)zirconium dimethyl,
bis(cyclopentadienyl)hafnium dimethyl,
bis(cyclopentadienyl)zirconium hydridochloride,
bis(cyclopentadienyl)hafnium hydridochloride,
15 bis(pentamethylcyclopentadienyl)zirconium dichloride,
bis(pentamethylcyclopentadienyl)hafnium dichloride, bis(n-
butylcyclopentadienyl)zirconium dichloride,
bis(dimethylcyclopentadienyl) zirconium dichloride,
cyclopentadienyl-zirconium trichloride,
20 bis(indenyl)zirconium dichloride, bis(4,5,6,7-tetrahydro-1-
indenyl)zirconium dichloride, and ethylene-[bis(4,5,6,7-
tetrahydro-1-indenyl)] zirconium dichloride. The
metallocene compounds utilized within the embodiment of
this art can be used as crystalline solids, as solutions in
25 aromatic hydrocarbons or in a supported form.

Multiple dry impregnations or carrier infusion steps
can be undertaken. That is, the step of alumoxane
impregnation of the carrier, without slurry formation, can
be undertaken, drying the pores under the conditions
30 specified above can be undertaken thereafter; these two
stages can be repeated one or more times to increase the
alumoxane or alumoxane/metallocene content of the catalyst.
Multiple infusion/drying steps are undertaken so that
relatively higher levels of alumoxane may be impregnated
35 within the silica pores without forming a layer of an
alumoxane coating on the external surface of the carrier.

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In the second infusion, dry impregnation, a solution can contain an alumoxane alone in solution or a solution of alumoxane and metallocene compound together.

As stated above, the alumoxane can be impregnated into the carrier at any stage of the process of catalyst preparation. When the catalyst contains two transition metals components, one of which is a metallocene, and one of which is a non-metallocene compound (free of unsubstituted or substituted cyclopentadienyl groups) the impregnation of the alumoxane in accordance with the unique method described above is preferably undertaken after hydroxyl groups of the carrier material are reacted with an organomagnesium compound and the non-metallocene transition metal compound. In this embodiment, the amount of Al, provided by alumoxane, is sufficient to provide an Al:transition metal (provided by metallocene) mole ratio ranging from 50 to 500, preferably 75 to 300. The carrier material, having said (OH) groups, is slurried in a non-polar solvent and the resulting slurry is contacted with at least one organomagnesium composition having the empirical formula below. The slurry of the carrier material in the solvent is prepared by introducing the carrier into the solvent, preferably while stirring, and heating the mixture to 25 to 70°C, preferably to 40 to 60°C. Temperatures here are critical with respect to the non-metallocene transition metal which is subsequently added; that is temperatures in this slurry of above 90°C result in deactivation of the non-metallocene transition metal added subsequently. The slurry is then contacted with the aforementioned organomagnesium composition, while the heating is continued at the aforementioned temperature.

The organomagnesium composition has the empirical formula $R_m Mg R'_n$, where R and R' are the same or different C_2-C_{12} alkyl groups, preferably C_4-C_{10} alkyl groups, more preferably C_4-C_8 alkyl groups, and most preferably both R and R' are mostly butyl groups, and m and n are each 0, 1

or 2, providing that $m + n$ is equal to the valence of Mg.

Suitable non-polar solvents are materials in which all of the reactants used herein, i.e., the organomagnesium composition, and the transition metal compound, are at least partially soluble and which are liquid at reaction temperatures. Preferred non-polar solvents are alkanes, such as isopentane, isohexane, n-heptane, octane, nonane, and decane, although a variety of other materials including cycloalkanes, such as cyclohexane, aromatics, such as benzene, toluene and ethylbenzene, may also be employed. The most preferred non-polar solvent is isopentane. Prior to use, the non-polar solvent should be purified, such as by percolation through silica gel and/or molecular sieves, to remove traces of water, oxygen, polar compounds, and other materials capable of adversely affecting catalyst activity.

In the most preferred embodiment of the synthesis of this catalyst it is important to add only such an amount of the organomagnesium composition that will be deposited - physically or chemically - onto the support since any excess of the organomagnesium composition in the solution may react with other synthesis chemicals and precipitate outside of the support. The carrier drying temperature affects the number of sites on the carrier available for the organomagnesium composition - the higher the drying temperature the lower the number of sites. Thus, the exact molar ratio of the organomagnesium composition to the hydroxyl groups on the carrier will vary and must be determined on a case-by-case basis to assure that only so much of the organomagnesium composition is added to the solution as will be deposited onto the support without leaving any excess of the organomagnesium composition in the solution. Furthermore, it is believed that the molar amount of the organomagnesium composition deposited onto the support is greater than the molar content of the hydroxyl groups on the support. Thus, the molar ratios

given below are intended only as an approximate guideline and the exact amount of the organomagnesium composition in this embodiment must be controlled by the functional limitation discussed above, i.e., it must not be greater than that which can be deposited onto the support. If greater than that amount is added to the solvent, the excess may react with the non-metallocene transition metal compound, thereby forming a precipitate outside of the support which is detrimental in the synthesis of our catalyst and must be avoided. The amount of the organomagnesium composition which is not greater than that deposited onto the support can be determined in any conventional manner, e.g., by adding the organomagnesium composition to the slurry of the carrier in the solvent, while stirring the slurry, until the organomagnesium composition is detected in the solvent.

For example, for the silica carrier heated at 600°C, the amount of the organomagnesium composition added to the slurry is such that the molar ratio of Mg to the hydroxyl groups (OH) on the solid carrier is 0.5:1 to 4:1, preferably 0.8:1 to 3:1, more preferably 0.9:1 to 2:1 and most preferably 1:1. The organomagnesium composition dissolves in the non-polar solvent to form a solution from which the organomagnesium composition is deposited onto the carrier.

It is also possible to add such an amount of the organomagnesium composition which is in excess of that which will be deposited onto the support, and then remove, e.g., by filtration and washing, any excess of the organomagnesium composition. However, this alternative is less desirable than the most preferred embodiment described above.

The organomagnesium treated support is contacted with an organic alcohol reagent (ROH) containing RO- groups which are reactive or capable of displacing alkyl groups on the magnesium. The amount of this organic alcohol reagent

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is effective to provide a ROH:Mg ratio of 0.5 to 2.0, preferably 0.8 to 1.5.

Contact of the silica supported magnesium compound, with the organic alcohol reagent is undertaken in the slurry. Contact is undertaken at a temperature ranging from 25°C to 80°C, preferably 40°C to 70°C.

The alkyl group in the organic alcohol reagent can contain 1 to 12 carbon atoms, preferably 1 to 8; in the embodiments below, it is an alkyl containing 2 to 4 carbon atoms, particularly of 4 carbon atoms (butyl). The inclusion of the alcohol reagent step in the catalyst synthesis of the invention produces a catalyst which, relative to the absence of this step, is much more active, requires much less non-transition metal (e.g. titanium), and is much more active with respect to the product component produced with the metallocene-transition metal component. After the addition of the organic alcohol reagent to the slurry is completed, the slurry is contacted with a non-metallocene transition metal compound, free of substituted or unsubstituted cyclopentadienyl groups.

The slurry temperature must be maintained at 25 to 70°C, preferably at 40 to 60°C. As noted above, temperatures in this slurry of 90°C or greater may result in deactivation of the non-metallocene transition metal. Suitable non-metallocene transition metal compounds used herein are compounds of metals of Groups IVA, and VA, of the Periodic Chart of the Elements, as published by the Fisher Scientific Company, Catalog No. 5-702-10, 1978, providing that such compounds are soluble in the non-polar solvents. Non-limiting examples of such compounds are titanium and vanadium halides, e.g., titanium tetrachloride, $TiCl_4$, vanadium tetrachloride, VCl_4 , vanadium oxytrichloride, $VOCl_3$, titanium and vanadium alkoxides, wherein the alkoxide moiety has a branched or unbranched alkyl radical of 1 to 20 carbon atoms, preferably 1 to 6 carbon atoms. The preferred non-metallocene transition

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metal compounds are titanium compounds, preferably tetravalent titanium compounds. The most preferred titanium compound is titanium tetrachloride. The amount of titanium or vanadium, in non-metallocene form ranges from a
5 Ti/Mg molar ratio of 0.3 to 1.5, preferably from 0.50 to 0.80.

Mixtures of such non-metallocene transition metal compounds may also be used and generally no restrictions are imposed on the transition metal compounds which may be
10 included. Any transition metal compound that may be used alone may also be used in conjunction with other transition metal compounds.

Incorporation of the alumoxane-metallocene can be directly to this slurry. Alternatively, and in accordance
15 with the unique method of infusion of alumoxane into the pores of the carrier, described above, the carrier slurry can be stripped of solvent, after the addition of the non-metallocene transition metal compound, to form a free-flowing powder. The free flowing powder can then be
20 impregnated by determining the pore volume of the carrier and providing an alumoxane (or metallocene-alumoxane) solution in a volume equal to or less than that of the pore volume of the carrier, and recovering a dry catalyst precursor. The resulting free-flowing powder, referred to
25 herein as a catalyst precursor, is combined with an activator (sometimes referred to as a cocatalyst). The cocatalyst can be a trialkylaluminum, free of alumoxane. Preferably, trimethylaluminum (TMA) is the cocatalyst or activator. The amount of the TMA activator is sufficient
30 to give an Al:Ti molar ratio of 10:1 to 1000:1, preferably 15:1 to 300:1, and most preferably 20:1 to 100:1. The catalyst exhibits high activity for long periods of time in the pilot plant, and exhibits little deactivation.

The catalyst precursor of this invention comprises a
35 metallocene compound and an alumoxane which is fed to the fluidized bed reactor for gas phase polymerizations and

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copolymerizations of ethylene in particulate form. Moreover, in accordance with the invention, the cocatalyst or activator is fed to the fluidized bed reactor for polymerizations and copolymerizations of ethylene in the absence of alumoxane solution as a cocatalyst.

EXAMPLES

EXAMPLE 1

Solution (A): To a 50 ml serum-bottle 0.140 grams of Cp_2ZrCl_2 was transferred and then 10.2 grams of a methylalumoxane (13.2 wt. % Al) solution were added. The solution remained at room temperature for 60 minutes until the entire contents were transferred to the silica slurry described below.

Into a 100 ml pear-flask equipped with a magnetic stirring bar, 3.0 grams of Davison 955 silica calcined at 600° , was added followed by addition of 20 ml dry toluene. The flask was placed into a 59°C oil bath. Next, 2.9 ml of dibutylmagnesium (0.74 mmol/ml) was added to the silica/toluene slurry. The contents of the flask were stirred for 25 minutes. Then, 2.3 mls of a 0.94 molar titanium tetrachloride solution in heptane was added to the flask. The slurry turned a dark brown color and stirring was continued for 25 minutes. Finally, the entire contents of solution (A) was transferred into the catalyst preparation flask, and the slurry was allowed to stir for 10 minutes. After this time, all solvents were removed by evaporation under a nitrogen purge. Catalyst yield was 5.6 grams of a dark-brown free-flowing powder. The Al/Zr ratio was 104.

EXAMPLE 2

Ethylene/1-hexene copolymer was prepared with the catalyst of the foregoing example under polymerization conditions to produce high density polyethylene (HDPE), with a flow index (I_{21}) of 6.

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A 1.6 liter stainless steel autoclave, at 50°C, was filled with 0.750 liters of dry heptane, 0.030 liters of dry 1-hexene and 4.0 mmols of trimethylaluminum (TMA) while under a slow nitrogen purge. The reactor was closed, the stirring rate was set at 900 rpm, the internal temperature was increased to 85°C, and the internal pressure was raised from 7 psi to 10 psi with hydrogen. Ethylene was introduced to maintain the reactor pressure at 203 psi. Next, 0.0639 grams of catalyst was introduced into the reactor with ethylene over pressure and the temperature was increased and held at 95°C. The polymerization was continued for 60 minutes, and then the ethylene supply was stopped and the reactor allowed to cool to room temperature. 78 grams of polyethylene were collected.

The molecular weight distribution (MWD) of the polymer was examined by Gel Permeation Chromatography (GPC), and the results clearly show that the polymer has a bimodal MWD.

Presently, commercial samples of HDPE with a bimodal MWD are produced in a tandem reactor process. In that process, two reactors are run in series and the catalyst is exposed to ethylene polymerization conditions in one reactor, and the resulting polymer-catalyst particles are transferred to a second reactor for additional polymerization. One of the main process differences in the two different reactors, is that the amount of hydrogen is different in the two different reactors. Relatively lower molecular weight product is produced in the reactor containing more hydrogen, because the hydrogen acts as a chain transfer agent; whereas relatively higher molecular weight product is produced in the reactor containing lesser relative amounts of hydrogen.

EXAMPLE 3

This catalyst was prepared in two stages. 495 grams of Davison grade 955 silica, previously calcined with dry

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nitrogen for 12 hours at 600°C, was added to a 2 gallon stainless steel autoclave under a slow nitrogen purge to eliminate oxygen and moisture from the catalyst preparation vessel. Then, 4.0 liters of dry isopentane (IC5) was added to the autoclave and the silica/IC5 were slurried at 100 rpm and the internal temperature was maintained at 55-60°C. Next, 469 ml of a 0.76 molar solution of dibutylmagnesium in heptane was added to the silica/IC5 slurry and stirring was continued for 60 minutes. Next, 39.1 ml of neat titanium tetrachloride was diluted with 40 ml of IC5 and this solution was added to the autoclave and stirring was continued for 60 minutes. Finally, the solvents were removed with a nitrogen purge through a vent line and 497 grams of a brown free-flowing powder were obtained. Ti found was 2.62 wt%; Mg found was 1.33 wt% and Ti/Mg molar ratio was 1.0.

492 grams of the product of the first stage was added to a 1.6 gallon glass catalyst preparation vessel fitted with a temperature jacket and an internal stirrer. Then into a stainless steel Hoke bomb was added 13.93 grams of (BuCp)₂ZrCl₂ (34.4 mmol Zr) and 717.5 ml of a methylalumoxane solution (3,444 mmol of Al) in toluene (4.8 Molar). Note: The total volume of the methylalumoxane/toluene solution is approximately equal to the total pore volume of the product of the first step. Next, the toluene solution containing the methylalumoxane and the zirconium compound were mixed and then the solution was added to the product of the first step in approximately 5 ml aliquots over 90 minutes; (during this time, the product of the first step remains completely dry and always consists of a free-flowing powder). Finally, nitrogen is purged through the glass vessel for 5 hours with the jacket temperature at 45°C. Yield: 877 grams of a free-flowing powder. Ti found was 1.85 wt%; Zr found was 0.30 wt%.

EXAMPLE 4

The catalyst described in Example 3 was examined in a pilot plant fluidized bed gas phase reactor under the following conditions:

5	ethylene	180 psi
	hydrogen/ethylene	0.005-0.008
	hexene/ethylene	0.015
	reactor temperature	95°C

The resin prepared at a productivity of 1400 g polymer/g catalyst had the following characteristics:

10	average particle size	0.017 inches
	resin metal content	13.0 ppm
	HLMI (I21)	5.3
	MFR (I21/I2.16)	113
15	Density	0.949 g/cc

Properties of films of the product of Example 4 are compared to the commercially produced product OxyChem L5005 commercially produced in a two stage reactor process, in which a different molecular weight component is made in each stage.

	Example 4	
Sample	<u>Ti/Zr</u>	<u>OxyChem L5005</u>
I ₂₁	5.3	8.0
MFR	113	160
25 Density	0.949	0.950
Throughput, lb/hr	98	120
Melt Pressure (at 120 lb/hr), psi	7550	6450
FQR	15	15
Dart Drop, 1 mil. g	565	325
30 0.5 mil, g	410	420
MD Elmend rf Tear, 0.5 mil. g/mil	37	25

The GPC results show that the Example 4 bimodal MWD product has a high molecular weight component with higher molecular weight than that produced in the tandem two reactor process. The film of Example 4 is substantially reduced in, if not free of, gel content. The film of the Example 4 product has improved dart impact strength.

EXAMPLE 5

Titanium Catalyst Component Preparation. 541 grams of Davison grade 955-[600°C calcination temperature] silica was weighed into a two-gallon stainless steel autoclave containing a stirring paddle. Next, ca. 4.8 liters of dry isopentane was added to the autoclave and the stirring rate was set at 100 rpm. The temperature of the silica/isopentane slurry was 54-58°C. Next, 546 mls of dibutylmagnesium (0.713 mmol/ml) was added to the slurry. The contents of the autoclave were stirred for 60 minutes. Then, 35.6 mls of neat 1-butanol were added and stirring was continued for one hour. Finally, 21.4 mls of titanium tetrachloride was added to the autoclave and stirring continued for 60 minutes. After this time, all solvents were removed by evaporation under a nitrogen purge. Catalyst yield was 496 grams of a tan free-flowing powder. Ti found 1.48 wt%; Mg found 1.48 wt.%.

EXAMPLE 6

Solution B: 12.95 grams of (BuCp)₂ZrCl₂ was transferred to a one-liter bottle and 687 ml of a 4.67 Molar (13.7 wt.% Al) solution of methylalumoxane were added. The bottle was shaken for one minute to form a yellow solution which was transferred into a 1.5 liter stainless steel hoke bomb and used immediately as described below.

Under an inert atmosphere, 465 grams of the titanium-containing catalyst described in Example 5 was added to a 2-gallon, glass-reactor vessel containing a helical stirrer to agitate the catalyst powder and a temperature jacket

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which was set at 30°C. The stirrer was set at 125 rpm. Then, the contents of the hoke bomb (solution B) was added to the titanium-containing catalyst in approximately 5-10 ml aliquots every 30-60 seconds over a 45 minute period.

5 The total volume of solution (B) used was such that the titanium containing catalyst always appeared "dry" during the entire addition time. However, during this addition time, the light tan titanium-containing catalyst turned a dark brown color. After the addition of solution (B) was
10 complete, the jacket temperature was set at 45°C and the residual toluene was removed with a nitrogen purge for 5 hrs. After this time the catalyst was a dark brown free-flowing powder. Analytical results: Mg, 0.99 wt.%; Ti, 0.96 wt.%; Al, 11.09 wt.% and Zr, 0.43 wt.%.

15 EXAMPLE 7

Polymerization - Slurry

Ethylene/1-hexene copolymer was prepared with the catalyst composition of Example 6 to produce HDPE with a relatively high molecular weight.

20 A 1.6 liter stainless steel autoclave, at 47°C. was filled with 0.750 liters of dry hexane, 0.030 liters of dry 1-hexene and 4.0 mmols of trimethylaluminum (TMA) was added while under a slow nitrogen purge. The reactor was closed, the stirring rate was set at 900 rpm, the internal
25 temperature was increased to 85°C. and the internal pressure was raised from 7 psi to 13 psi with hydrogen. Ethylene was introduced to maintain the reactor pressure at 200 psi. Next, 0.0244 grams of catalyst composition of Example 6 was introduced into the reactor with ethylene
30 over-pressure and the temperature was increased and held at 95°C. The polymerization was continued for 60 minutes, and then the ethylene supply was stopped and the reactor allowed to cool to room temperature. 36.6 grams of polyethyl n were collected. The HLMI of th polym r was
35 4.3 indicating a r latively high molecular weight.

EXAMPLE 8**Polymerization - Gas Phase Fluidized Bed Reactor**

The catalyst of Example 6 was added to a fluidized bed gas phase reactor under the following conditions:

5	Ethylene	185 psi
	H ₂ /C ₂	0.009
	Hexene/ethylene	0.013
	Reactor Temp.	90°C
10	Cocatalyst	TMA (trimethylaluminum) 300 ppm based on ethylene feed based on ethylene feed

EXAMPLE 9**Product - Alpine Film Line:**

15 Polymer produced in Example 8 with a HLMI of 5.9, MFR of 128, which gave a GPC chromatogram with a HMW/LMW ratio of 64.2/35.8, was fabricated into HDPE film at 0.5 and 1.0 mil. The impact data is shown below:

20	RESIN	FI/MFR(*)	1.0 mil Dart Drop	0.5 mil Dart Drop
	Example 4 herein	5/111	752	415
	Tandem Reactors	7/110	275	395

(*) values for pellets obtained after stabilizer package was added to granular resin

25 EXAMPLE 10

(A) - Titanium component: At ambient temperature, 509 grams of a Davison grade 955 silica, previously dried at 600°C, was added to a 2-gallon stainless steel catalyst preparation vessel under a slow nitrogen purge. Four
30 liters of dry isopentane was added to the vessel to form a silica/isopentane slurry and the stirring rate was adjusted to 100 rpm. The temperature of the slurry was set at 55°C

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and a solution of dibutylmagnesium in heptane (366 mmol Mg) was added and stirring was continued for 70 minutes. Then, 439 mmol of 1-butanol was added. After 65 minutes, 256 mmol of TiCl_4 was added and stirring continued for 60 minutes. Finally, the solvents were removed by evaporation with a nitrogen purge. 518 grams of a white, free-flowing powder were obtained. Analytical: Ti, 1.89 wt%; Mg, 1.36 wt%.

(B) - Zirconium component: 490 grams of the product prepared in Part A was added to a 2-gallon glass catalyst preparation vessel fitted with a temperature jacket and an internal stirrer. The entire contents of solution (B) described below was added to the part A product in 5-10 ml aliquots, which were added in 1 minute intervals. 1.79 ml of solution (B) was used per gram of the titanium component. Finally, the jacket temperature was set at 45°C and residual toluene was removed under a nitrogen purge. 642 grams of a dark brown powder were obtained. Analytical: Ti, 1.17 wt%; Mg, 0.85 wt%; Al, 13.5 wt%; and Zr, 0.43 wt%.

Solution (B) - Under an inert atmosphere, 16.85 grams of $(\text{BuCp})_2\text{ZrCl}_2$ was added to a 1-liter glass bottle followed by 877 mls of a 4.75 Molar (Al) methylalumoxane solution (in toluene). The bottle was closed and shaken vigorously for one minute to dissolve all solids. The contents of the flask were transferred to a 1.5 liter stainless steel Hoke bomb and added to the part A product as described above.

EXAMPLE 11

(A) - Titanium component: At ambient temperature, 500 grams of a Davison grade 955 silica, previously dried at 600°C, was added to a 2-gallon stainless steel catalyst preparation vessel under a slow nitrogen purge. Four liters of dry isopentane was added to the vessel to form a silica/isopentane slurry and the stirring rate was adjusted to 100 rpm. The temperature of the slurry was set at 55°C

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and a solution of dibutylmagnesium in heptane (352 mmol Mg) was added and stirring was continued for 60 minutes. Next, 359 mmol TiCl_4 was added and stirring continued for 60 minutes. Finally, the solvents were removed by evaporation with a nitrogen purge. 604 grams of a white, free-flowing powder were obtained. Analytical: Ti, 3.03 wt%; Mg, 1.54 wt%.

(B) - Zirconium component: 378 grams of the product prepared in Part A was added to a 2 gallon glass catalyst preparation vessel fitted with a temperature jacket and an internal stirrer. The entire contents of solution (B) (described below) was added to the part A product in 5-10 ml aliquots. The aliquots were added in 1 minute intervals. 2.03 ml of solution (B) was used per gram of the titanium component. Finally, the jacket temperature was set at 45°C and residual toluene was removed under a nitrogen purge. 498 grams of a dark brown powder were obtained. Analytical: Ti, 1.42 wt%; Mg, 0.78 wt%; Al, 11.9 wt%; and Zr, 0.43 wt%.

Solution (B) - Under an inert atmosphere, 14.5 grams of $(\text{BuCp})_2\text{ZrCl}_2$ was added to a 1-liter glass bottle followed by 769 mls of a 4.67 Molar (Al) methylalumoxane solution (in toluene). The bottle was closed and shaken vigorously for one minute to dissolve all solids. The contents of the flask were transferred to a 1.5 liter stainless steel Hoke bomb and added to the part A product as described above.

EXAMPLE 12

Raw materials used in catalyst preparation included 504 g of Davison grade 952-1836 silica, 677 g of methylalumoxane in toluene solution (30 wt.% MAO), 7.136 g of bis(n-butylcyclopentadienyl)zirconium dichloride.

The steps of the catalyst preparation are set forth below:

1. Dehydrate the 952-1836 silica at 250°C for 4 hours using air to purge. Then purge with nitrogen on cooling.
2. Transfer the silica to a mix-vessel.
- 5 3. Add 7.136 g of bis(n-butylcyclopentadienyl) zirconium dichloride and 677 g of methylalumoxane to a bottle.
4. Agitate the catalyst solution in the bottle until the metallocene dissolves in the MAO solution.
- 10 5. Transfer the MAO and metallocene solution into the mix-vessel containing the dehydrated 952-1836 silica slowly while agitating the silica bed vigorously to make sure that the catalyst solution is well dispersed into the silica bed.
- 15 6. After the addition, continue to agitate the catalyst for 1/2 hours.
7. Start drying the catalyst by purging with nitrogen for 5 hours at 45°C.
- 20 8. Sieve the catalyst to remove particles larger than 150 micron.
9. The catalyst has the following analysis:
Yield = 767 g catalyst (from 500 g of silica)
Al = 9.95 wt. %
Zr = 0.19 wt. %

25 EXAMPLE 13

Produce a polymer for low density film, 0.918 g/cc, 1.0 MI, 17 MFR, in a fluidized bed, gas phase reactor using the following process conditions:

Process Conditions:

	Fluidization velocity	1.7 ft/sec
	Residence time	2.5 hours
	Temperature	84°C
5	Ethylene	220 psi
	Hexene	3.6 psi
	Isopentane	50 psi
	Carbon dioxide	3 ppm
	Ash	200 to 250 ppm
10	The catalyst was that of Example 12.	

COMPARATIVE EXAMPLESComparative Example 1

A zirconium catalyst was tested in a slurry reactor at 85°C with 130 psi ethylene partial pressure. A
15 hexene/ethylene gas ratio of 0.03 was used. MAO/toluene solution (12 wt.%, 2 ml) was added to the reactor. Productivity of 800 g resin/g catalyst/hr was measured.

The same catalyst system was tested in the fluidized bed reactor at 90°C with 200 psi ethylene partial pressure.
20 A 0.025 hexene to ethylene gas ratio was used. A feed rate of 150 to 200 cc/hr of 2 wt% MAO/toluene solution was employed. The MAO solution was added below the distributor plate. Even at very high MAO/toluene feed rates, catalyst productivity was only 220 g resin/g catalyst/hr. In
25 addition, the reactor had to be shut down due to a fouled plate only 18 hours after the MAO feed was started.

This example illustrates that it is more effective to activate zirconium catalysts prior to introduction into a gas phase reactor. It also illustrates the fouling
30 problems experienced when MAO solutions are added to the gas-phase reactor.

Comparative Example 2

A titanium/zirconium mixed metal catalyst was tested in the fluidized bed reactor. At 150 psi ethylene partial

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pressure at 90°C a 0.04 hexene to ethylene gas ratio was employed, and a hydrogen to ethylene gas ratio was 0.045 . A 2 wt% solution of MAO in toluene was added beneath the fluidized bed distributor plate. Resin flow index and GPC curve analysis showed that the zirconium catalyst sites were active, and the Ti:Zr productivity ratio was 7:3. However, the reactor had to be shut down within 24 hours because the distributor plate had fouled.

Comparative Example 3

The same titanium/zirconium catalyst used in Comparative Example 2 was tested in the fluidized bed reactor. It was run at 90°C with 150 psi ethylene partial pressure. A 0.03 hexene to ethylene gas ratio was used and a hydrogen to ethylene ratio was 0.04 . A solution of 2 wt% MAO in toluene was added directly into the bed at the rate of 200 cc/hr. The resin flow index and molecular weight distribution showed definitively that the zirconium sites were active with a Ti:Zr productivity ratio of 3:7. In the process of running this test, though, a very large chunk grew around the injection port causing a shutdown.

This example demonstrates that relative zirconocene catalyst activity is significantly higher when there is better contact between the MAO/toluene droplets and the catalyst sites. It also verifies that fouling also occurs when the MAO solution is added to the reactor directly into the fluidized bed of polymer.

Comparative Example 4

The catalyst used in Comparative Examples 2 and 3 was re-run under the same conditions used in Comparative Example 3. The MAO feed rate was the same as well. During this test, though, the MAO was dispersed into a 10 lb/hr ethylene gas stream using an ultrasonic atomizer. The atomizer disperses the MAO solution into very small (40 micron) droplets.

Enough gas was used so that the toluene evaporated from the MAO. The gas flow rate was determined in an off-

line test using toluene alone. The resin produced during this test showed no evidence of activity from the zirconium sites. In addition, there were no signs of reactor fouling after an extended period of running.

5 This example proves that it is the presence of liquid in the reactor that is responsible both for the activation of the zirconium and the fouling of the reactor.

10 Thus it is apparent that there has been provided, in accordance with the invention, a process that fully satisfies the objects, aims, and advantages set forth above. While the invention has been described in conjunction with specific embodiments thereof, it is evident that many alternatives, modifications, and variations will be apparent to those skilled in the art in
15 light of the foregoing description. Accordingly, it is intended to embrace all such alternatives, modifications, and variations as fall within the spirit and broad scope of the appended claims.

WHAT IS CLAIMED:

1. A process for supporting alumoxane and derivatives thereof on a support, without forming a slurry of the support during the deposition of alumoxane on the support, wherein the process comprises

5 1) providing a sample comprising a support wherein the sample has a total pore volume; and wherein the support has a pore volume of 0.5 to 5.0 cc/g and has pores of an average diameter of 50 to 500 angstroms;

10 2) providing a volume of a solution of an alumoxane and a solvent for said alumoxane, wherein the volume of said solution is in a range of from equal to the total pore volume of the sample and less than twice the total pore volume of the sample and;

15 wherein the volume of said solution is less than that required to form a slurry of said sample in said solution;

20 wherein the alumoxane is of formula (a) and/or(b), wherein (a) is $R-(Al(R)-O)_x-AlR_2$ for linear alumoxanes and (b) $(-Al(R)-O)_y$ for cyclic alumoxanes and wherein x is 1 to 40, y is 3-40 and R is a C_1-C_8 alkyl group

25 wherein the solution has a concentration of alumoxane expressed as Al weight percent which ranges from 5 to 20,

 wherein the alumoxane provides aluminum in an amount sufficient to provide a ratio of Al to silica (wt/wt) from 0.10 to 0.40;

30 (3) contacting the sample with said volume of said solution and allowing the solution to impregnate the pores of support, without forming a slurry of the support in the solvent;

 (4) after said contacting recovering particles of support impregnated with alumoxane.

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2. The process of Claim 1, which further comprises heating the dry particles to remove solvent from the pores under temperature conditions effective to prevent degradation of the alumoxane.

3. The process of Claim 2, wherein the temperature ranges from above 30° to below 60°C.

4. The process of Claim 1, wherein the support is silica.

5. The process of Claim 4, wherein the solution further comprises a metallocene compound.

6. The process of Claim 5, which further comprises adding, to said volume of solution prior to said contacting, at least one metallocene compound of the formula: $Cp_mMA_nB_p$, wherein

5 Cp is a cyclopentadienyl or a substituted cyclopentadienyl group;

m is 1 or 2;

M is zirconium or hafnium; and

10 each of A and B is selected from the group consisting of a halogen atom, a hydrogen atom and an alkyl group, providing that $m+n+p$ is equal to the valence of the metal M; wherein the metallocene compound is admixed with an amount of a methylalumoxane effective to activate the metallocene compound to form a catalyst.

7. The process of Claim 6, wherein the metallocene compound is selected from the group consisting of bis(cyclopentadienyl)metal dihalides, bis(cyclopentadienyl)metal hydridohalides, 5 bis(cyclopentadienyl)metal monoalkyl monohalides, bis(cyclopentadienyl)metal dialkyls and bis(indenyl)metal dihalides wherein the metal is zirconium or hafnium, halide groups are preferably chlorine and the alkyl groups are C₁-C₆ alkyls.

8. The process of Claim 6, wherein the metallocene compound is selected from the group consisting of bis(cyclopentadienyl)zirconium dichloride, bis(cyclopentadienyl)hafnium dichloride, 5 bis(cyclopentadienyl)zirconium dimethyl, bis(cyclopentadienyl)hafnium dimethyl, bis(cyclopentadienyl)zirconium hydrido-chloride, bis(cyclopentadienyl)hafnium hydrido-chloride, bis(pentamethylcyclopentadienyl)zirconium dichloride, 10 bis(pentamethylcyclopentadienyl)hafnium dichloride, bis(n-butylcyclopentadienyl)zirconium dichloride, bis(dimethylcyclopentadienyl) zirconium dichloride, cyclopentadienyl-zirconium trichloride, bis(indenyl)zirconium dichloride, bis(4,5,6,7-tetrahydro-1- 15 indenyl)zirconium dichloride, and ethylene-[bis(4,5,6,7-tetrahydro-1-indenyl)] zirconium dichloride.

9. The process of Claim 8, wherein the solution has a composition which provides a molar ratio of alumoxane, expressed as aluminum, to metallocene ranging from 50 to 500.

10. The process of Claim 9, wherein the alumoxane is methylalumoxane.

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11. The process of Claim 10, which further comprises heating the dry particles to remove solvent from the pores under temperature conditions effective to prevent degradation of the alumoxane.

12. The process of Claim 11, wherein the temperature ranges from above 30° to below 60°C.

13. The process of Claim 4, wherein the sample further comprises TiCl_4 , wherein the TiCl_4 is deposited in said pores.

14. The process of Claim 13, wherein the sample further comprises a source of magnesium.

15. The process of Claim 14, wherein the sample is formed by contacting the silica with a dialkylmagnesium, in which the alkyl contains 1 to 8 carbon atoms to form a first intermediate; contacting the first intermediate with
5 an alcohol $[\text{R}''\text{OH}]$ in which the $\text{R}''\text{OH}/\text{Mg}$ ratio is 0.50 to 2.0 in which R'' is alkyl of 1 to 10 carbon atoms to form a second intermediate.

16. The process of Claim 15, wherein the second intermediate is contacted with TiCl_4 .

17. The process of Claim 16, wherein the sample is dry and wherein it is formed by providing silica which contains reactive hydroxyl groups; contacting said hydroxyl groups on said silica with an amount of $R_aMgR'_b$,

5 wherein R and R' are the same or different C_2-C_8 alkyl groups, and a and b are each 0, 1, or 2, providing that $a + b$ is equal to the valence of Mg, and

10 wherein the amount of $R_aMgR'_b$ is sufficient to provide a Mg:OH molar ratio which ranges from 0.5:1 to 4:1; and

15 adding butanol to said $R_aMgR'_b$ contacted hydroxyl groups and thereafter introducing said non metallocene transition metal compound to said second intermediate recovering dry particles for use as said sample.

18. The process of Claim 17, wherein the solution further comprises a metallocene compound.

INTERNATIONAL SEARCH REPORT

International application N .
PCT/US95/12743**A. CLASSIFICATION OF SUBJECT MATTER**

IPC(6) : B01J 31/00, 37/00; C08F 4/00, 4/42, 4/646

US CL : 502/103, 104, 107, 110, 113, 115, 117, 120

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

U.S. : 502/103, 104, 107, 110, 113, 115, 117, 120

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched
NONEElectronic data base consulted during the international search (name of data base and, where practicable, search terms used)
NONE**C. DOCUMENTS CONSIDERED TO BE RELEVANT**

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
Y	US,A 5,026,797 (TAKAHASHI) 25 JUNE 1991, COLUMN 3, LINES 4-10 AND 40-58, COLUMN 6, LINES 56-63	ALL
Y	US,A 5,032,562 (LO ET AL.) 16 JULY 1991, COLUMN 3, LINES 44-52 AND 60-68, COLUMN 4, LINES 16-45 AND COLUMN 5, LINES 1-49	ALL
Y	US,A 5,001,205 (HOEL) 19 MARCH 1991, COLUMN 8, LINES 44-50 AND COLUMN 9, LINES 3-10	ALL
Y	US,A 5,240,894 (BURKHARDT ET AL.) 31 AUGUST 1993, COLUMN 6, LINES 41-66	ALL
X	US,A 5,332,706 (NOWLIN ET AL.) 26 JULY 1994, COLUMN 15, LINE 40-COLUMN 17, LINE 32	ALL

☐ Further documents are listed in the continuation of Box C. ☐ See patent family annex.

* Special categories of cited documents:	"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
"A" document defining the general state of the art which is not considered to be of particular relevance	"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
"B" earlier document published on or after the international filing date	"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art
"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)	"A" document member of the same patent family
"O" document referring to an oral disclosure, use, exhibition or other means	
"F" document published prior to the international filing date but later than the priority date claimed	

Date of the actual completion of the international search

09 FEBRUARY 1996

Date of mailing of the international search report

21 FEB 1996

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